

SUL ROSS STATE UNIVERSITY GENERAL CHEMISTRY II

1112—LAB SYLLABUS—Spring 2016

LAB: General Chemistry 1112
Room: WSB 307
Time: L01 T 1:00-2:50 pm
L02 R: 3:00 pm-4:50 pm

TA: Tres Wright
Office: WSB 308
Email: francis.wright@sulross.edu
Hours: Mon: 1–4 Tues: 3–5 Wed: 1–2
Thurs: 1–5 Fri: 1–4

Required Laboratory Manual

General Chemistry II: Laboratory Manual CHEM 1112 by Dr. Leaver, 2015 (available from the bookstore)

OBJECTIVES

Program Learning Objectives (PLO):

A student graduating with the ***chemistry major*** is expected to demonstrate that (s)he is able to:

1. Explain atomic & molecular structures, bonding, thermodynamics, chemical equilibria & kinetics, stoichiometry, and electrochemical processes
2. Write and explain organic reactions, stereochemistry, and reactions in biological systems
3. Use essential modern instruments to perform chemistry experiments in the laboratory
4. Summarize basic principles of research design and analyze experimental data using appropriate computer programs (e.g. Excel, Sigma-plot, etc.) in regards to the chemistry discipline.

Core Objectives (CO):

A. Critical Thinking Skills – Students will gain/improve their critical thinking ability by solving real life chemistry problems through inquiry, analysis, and evaluation of available information. Students will be tested on their critical thinking ability in exams and through lab experiments

B. Communication Skills – Students will have the opportunity to improve communication skills through oral discussion and writing reports (i.e. observation, explanation, and conclusion etc.) on the experiments done in the lab sessions.

C. Empirical and Quantitative Skills – students will use the mathematical skills needed to manipulate and analyze numerical data obtained through experimentation in order to form conclusions

D. Teamwork – students will use team-spirit and consider different points of view to work effectively while conducting experiments as a team working toward a shared purpose or goal

Expectations:

- Read over the experiment before lab
- Follow all safety procedures:
- Shorts, flip-flops, chewing gum and open-toed shoes are **NOT** allowed in lab. If you come to class without appropriate clothing, you will be asked to leave.
NO EXCEPTIONS!

Attendance:

Coming to lab is mandatory. Be on time and SIGN IN at the beginning of the lab period. Plan to spend the entire period in lab. The TA may deduct points for students who arrive late or leave early.

If you *miss 3 labs* or more you will receive an automatic *F*.

Assignments:

Lab Manual Assignments:

- Pre-Lab: Due at the beginning of the lab that the experiment will be performed
- Data Sheet: Due the lab period after the lab is done

Written Assignments:

- Pre-Lab Write Up: Due at the beginning of the lab that the experiment is performed
- Lab Report: Due the lab period after the lab is done
- Guidelines for writing lab reports are shown on the next page and are found in your General Chemistry II Laboratory Manual

Lab Grading:

- Each experiment is worth 30 points
- These points will come from:
 - o The pre-lab definitions (5)
 - o The experiment work in the manual (25)
 - o Points will be deducted for not turning in lab report and write up
- Assignments must be completed and turned in on time
 - o Assignments must be legible
 - o Assignments and reports will be due the following week unless told otherwise

- o 10% of the grade will be deducted for assignments not turned in at the beginning of lab. An additional 10% will be deducted for each day that the assignment is late

Outline for Written Pre-lab Reports

- **Aim:** Here you will state the goal of the experiment (in your own words).
- **Reagents:** You will make a list of all of the chemicals used in the experiment along with relevant data (grams, volume, molarity etc. that is indicated in your experiment).
- **Apparatus:** You will list all of the equipment that you will use.
- **Method:** This is where you will outline the steps in the experiment. The steps will be put in your own words.

Outline for Written Lab Reports (Use Microsoft Word or related program)

- **Aim:** Here you will state the goal of the experiment (in your own words).
- **Reagents:** You will make a list of all of the chemicals used in the experiment along with relevant data (grams, volume, molarity etc.). This is how much **YOU** used, not how much the manual asks for.
- **Apparatus:** You will list all of the equipment that you used.
- **Method:** This is where you will outline the steps in the experiment. Be sure to note any difference between what you did and what the manual said to do.
- **Data & Results:** Note the observations that you made during the experiment. What are your findings? (percent yield, melting point, etc.)
- **Discussion:** Discuss your results and answer the questions that were asked in the **Data and Results** section of the experiment. Talk about

the significance of your results. Were your results expected or unexpected? Why or why not?

- **Conclusion(s):** Summarize the key points and findings of the experiment. Was the experiment successful or unsuccessful?

Students with Special Needs: *Sul Ross State University is committed to equal access in compliance with the Americans With Disabilities Act of 1973. It is the student's responsibility to initiate a request for accessibility services. Students seeking accessibility services must contact Mary Schwartz in Counseling and Accessibility Services, Ferguson Hall, Room 112. The mailing address is P.O. Box C-122, Sul Ross State University, Alpine, Texas 79832. Telephone: 432-837-8203. E-mail: mschwartz@sulross.edu.*

Scholastic Dishonesty: Students who violate the University rules on scholastic dishonesty are subject to penalties, including the possibility of an **F** in the course and/or dismissal from the University.

General Chemistry 1112 Lab Schedule

<u>Week of</u>	<i>EXPERIMENT</i>
January 18	<i>No Laboratory Experiment</i>
January 25	Safety Practices and Procedures in the Laboratory (Exp.1) & syllabus discussion
February 1	Heat of Reaction (Exp. 2)
February 8	Enthalpy of Solutions (Exp. 3)
February 15	Polymers: A Ticky-Tacky Plastic World (Exp. 4)
February 22	Colligative Properties (Exp. 5)
February 29	Rates of Chemical Reactions (Exp. 6)
March 7	Acid base Titration (Exp. 7)
March 14-18th	<i>Spring Break (No classes)</i>
March 21	Determination of Dissociation Constant of a Weak Acid (Exp. 8)
March 28	Titration of Polyprotic Acids (Exp. 9)
April 4	Solubility of Practically Insoluble Salts (Exp. 10)
April 11	Qualitative Analysis of Cations (Exp. 11)
April 18	Qualitative Analysis of Anions (Exp. 12)
April 25	Lab Clean Up and Final